

PHARMACEUTICAL ANALYSIS – BP102T

UNIT: III PRECIPITATION TITRATION, COMPLEXOMETRIC TITRATION,
GRAVIMETRY AND DIAZOTISATION TITRATION

CLASS:23

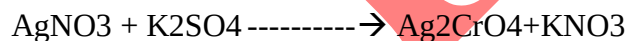
TOPIC: ESTIMATION OF SODIUM CHLORIDE**ESTIMATION OF 1.0M NaCl [MOHR'S METHOD] :**

Preparation of 0.1M AgNO₃:

Take 10 gm of silver nitrate, which has previously been dried at 120°C for 2 hrs, and cool it in a desiccator. Then, take 8.49 gm of AgNO₃ and mix it in sufficient water to produce 500 ml, which is equivalent to 0.1 M.

Preparation of 0.1M NaCl:

Take 2.922 gm of sodium chloride and mix in sufficient water to produce 500ml which is equivalent to 0.1M NaCl

Principle:**Method:**

- Take 10 ml of 0.1M NaCl in a conical flask.
- Add a few drops of potassium chromate as an indicator.
- The Burette is filled with 0.1M AgNO₃. During the titration, silver nitrate reacts with sodium chloride and forms a white ppt of AgCl.
- Silver nitrate reacts with the indicator and forms silver chromate which gives brick

red colour indicating the endpoint

- Report the process three times to get an accurate endpoint.
- Take the average value, go for necessary calculation
- Each ml 0.1M AgNO_3 = 0.005845 gm of NaCl

RCR