

Eutectic mixtures :

- Eutectic mixtures are defined as mixtures of low melting point ingredients which on mixing together turn to liquid form due to depression in melting point of the mixture below room temperature
- They are mixture of substances, that liquify when mixed, rubbed or triturated together.
- The melting point of many eutectic mixtures

- are below room temperature
- Examples of the substances that liquify when mixed or triturated together
- The melting points of many Eutectic mixtures are below room temperature.
- Examples of the substances which tend to liquify on mixing are Camphore, Thymol, menthol & Salol
- Any two of these drugs turn to liquid when mixed.
- This problem during formulation of powders of such material can be solved by using inert adsorbent such as talc, starch, lactose to prevent dampness of the powder and dispersing the components of the Eutectic mixture separately.

Geometric Dilutions::

- Geometric Dilution is the process by which homogeneous mixture or even distribution of 2 or more substance is achieved.
- This method is used when potent substances must be mixed with a large amount of diluent

- The potent drug and an approximately equal volume of diluent are placed in a mortar and thoroughly mixed by trituration.
- A second portion of diluent equal in volume to the powder mixture in the mortar is added, and trituration is repeated.
- The process is continued equal volumes of diluent are added to the powder mixture in the mortar until all of the diluent is incorporated.

Ex:- Dose of potent drug is 120 mg

- potent drug (120)mg + 120 mg of diluent
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are added and mixed thoroughly.
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240 mg of potent drug & diluent are
again mixed with further 240 mg of
diluent
↓

This process is repeated until all the
diluent is incorporated.